

Review Stoichiometry Section 2 Answers With Work

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Review Stoichiometry Section 2 Answers

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N₂ are mixed with 120 mol of H

Chapter 9 Review Stoichiometry Section 2 Answers Modern ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed?

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

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Pg 311 Section Review Stoichiometry Answers

$2\text{O}_3(\text{l}) \rightarrow 4\text{Al}(\text{s}) + 3\text{O}_2(\text{g})$ Recall from your study of equations and reactions that the coefficients in a chemical equation satisfy the law of conservation of matter and represent the relative amounts in moles of reactants and products. Therefore, 2 mol of aluminum oxide decomposes to produce 4 mol of aluminum and 3 mol of oxygen gas.

CorrectionKey=NL-A DO NOT EDIT--Changes must be made ...

$2\text{O} = 204.1 \text{ g products}$ Add the masses of the products. $204.1 \text{ g reactants} = 204.1 \text{ g products}$ The law of conservation of mass is observed. 3 Evaluate the Answer. The sums of the reactants and the products are correctly stated to the first decimal place because each mass is accurate to the first decimal place.

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Answers With Work

Chapter 11: Stoichiometry

SECTION 2 continued Date Class _____ 60.2 9 42.1 1 a. \ tt mash
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density of 1.43 g/L calculate the volume of this gas. 76
STOICHIOMETRY MODERN CHEMISTRY a. —. 81 g 6. A car air bag
requires 70. L of nitrogen gas ...

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